

# FIFTY FREQUENTLY FORGOTTEN FUN FACTS

This packet contains topics from each of the units we worked on this year with questions. Most of the questions are similar to what you would expect to see on Part B2 and C of the Regents Exam in Chemistry. The multiple choice questions mirror common questions found on Parts A and B1.

## I. ATOMIC STRUCTURE & NUCLEAR CHEMISTRY

1) Protons are +1 each with a mass of 1 amu each, the number of protons = atomic number, nuclear charge = + (# protons). [Periodic Table]

- How many protons are there in a nucleus of Kr-85? \_\_\_\_\_
- What is the nuclear charge of an atom of Br? \_\_\_\_\_
- What is the mass of the protons in a nucleus of O-15? \_\_\_\_\_

2) Neutrons are neutral with a mass of 1 amu each, # neutrons = mass # - atomic number. Isotopes = atoms of the same element (same atomic #) but different # of neutrons (mass #). [Periodic Table]

- How many neutrons are there in the nucleus of  $^{56}_{26}\text{Fe}$ ? \_\_\_\_\_
- Circle the two nuclei that are isotopes of each other:  $^{16}_8\text{O}$      $^{14}_7\text{N}$      $^{16}_8\text{O}$      $^{19}_9\text{F}$

3) Electrons are each -1 with a mass that is VERY, VERY tiny compared to the mass of a proton or neutron.

a) Which particle has a mass that is 1/1836<sup>th</sup> the mass of a proton?

- $^4_2\text{He}$
- $^1_1\text{H}$
- $^{0}_{-1}\text{e}$
- $^1_0\text{n}$

4) Natural Decay: Parent Nuclide → Decay particle + daughter nuclide [Tables N and O]

- Write the decay for U-238: \_\_\_\_\_
- Write the decay for K-37: \_\_\_\_\_
- Write the decay for P-32: \_\_\_\_\_

5) Artificial Transmutation is when a relatively stable nucleus is impacted by a particle bullet at high speeds and becomes an unstable nucleus of a different element. Nuclear fission occurs when nuclei of U-235 or Pu-239 are impacted by a neutron and split into two smaller nuclei and more neutrons. Nuclear fusion occurs when two small nuclei of hydrogen combine at high temperatures and pressures to form larger nuclei of helium. Both fission and fusion convert mass into a huge amount of energy.

Given the nuclear reactions:

- $^{235}_{92}\text{U} + ^1_0\text{n} \rightarrow ^{92}_{36}\text{Kr} + ^{141}_{56}\text{Ba} + 3 ^1_0\text{n}$
- $^{239}_{94}\text{Pu} + ^4_2\text{He} \rightarrow ^{242}_{96}\text{Cm} + ^1_0\text{n}$
- $^{226}_{88}\text{Ra} \rightarrow ^{222}_{86}\text{Rn} + ^4_2\text{He}$
- $^2_1\text{H} + ^2_1\text{H} \rightarrow ^4_2\text{He}$

- Which reaction represents natural decay? \_\_\_\_\_
- Which reaction represents artificial transmutation? \_\_\_\_\_
- Which reaction represents nuclear fission? \_\_\_\_\_
- Which reaction represents nuclear fusion? \_\_\_\_\_

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**United States. Congress**



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